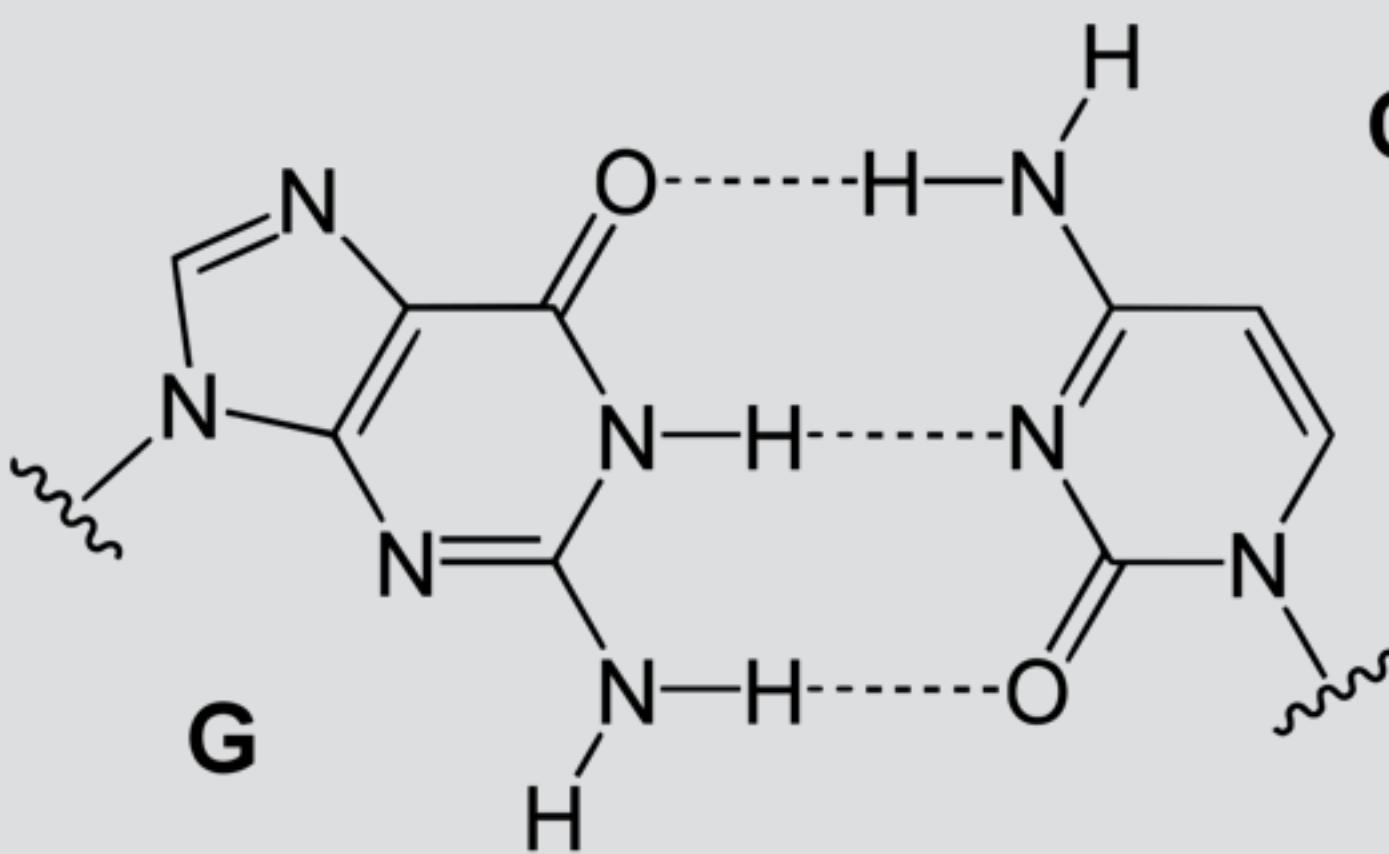
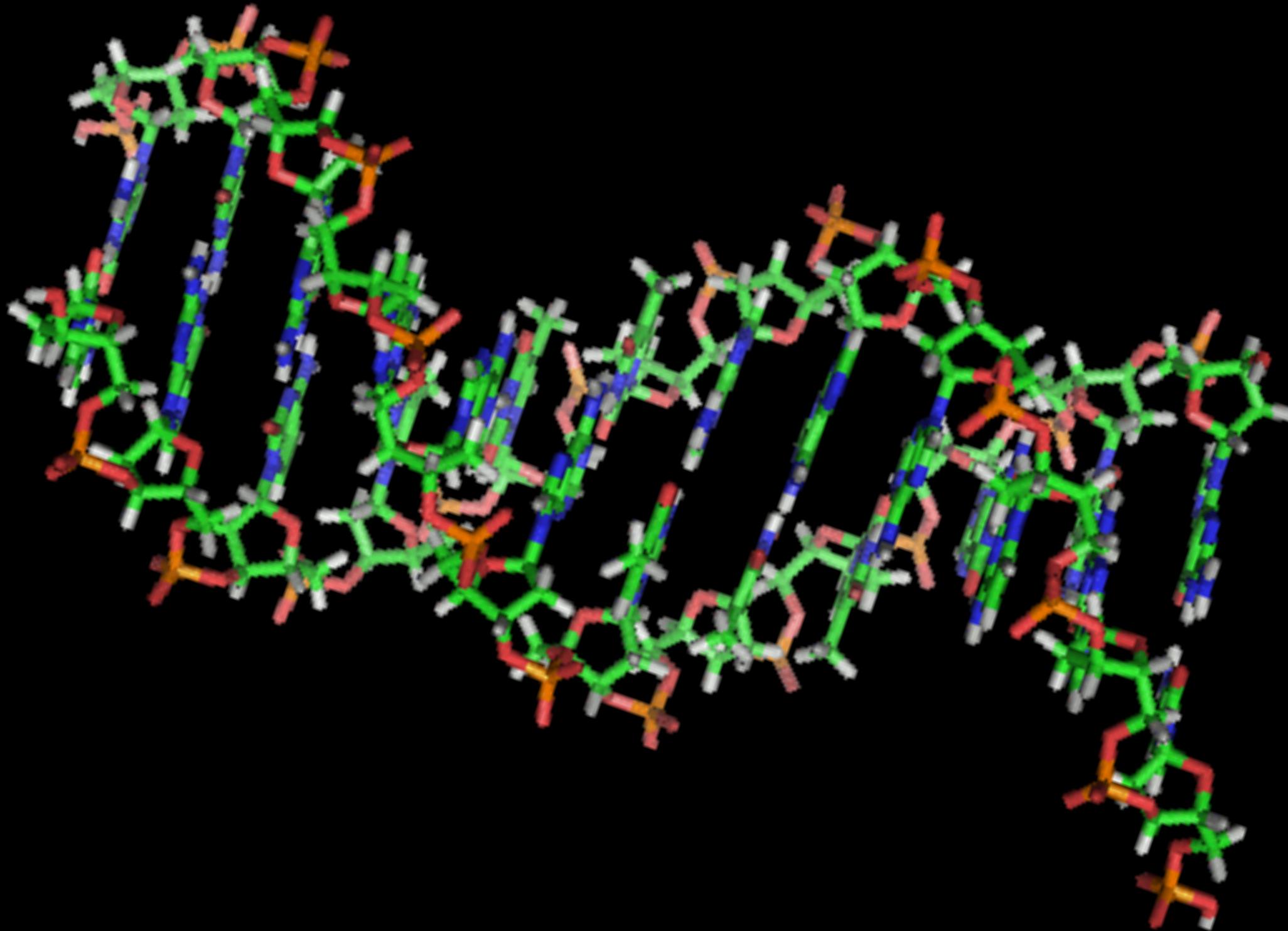
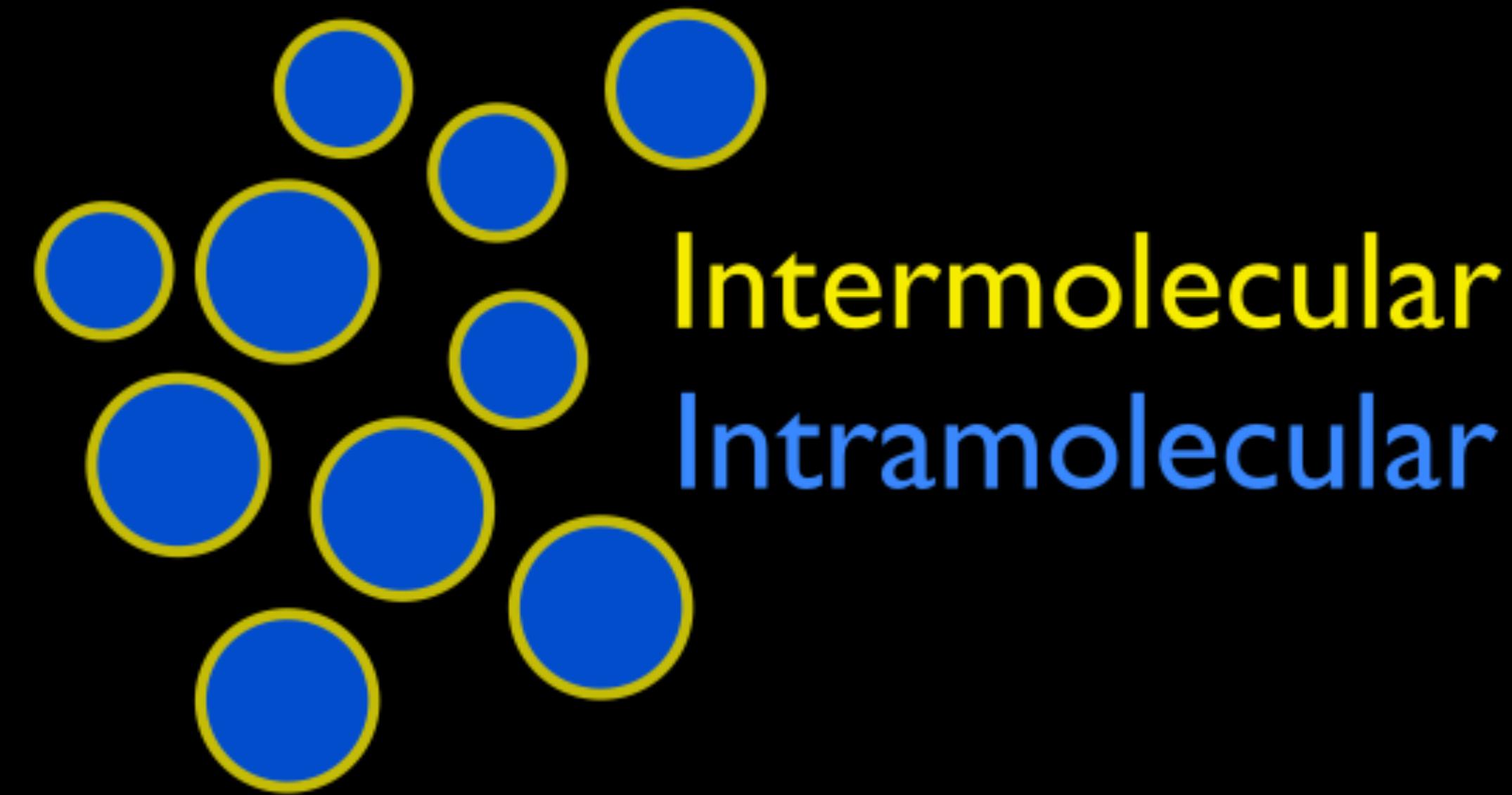


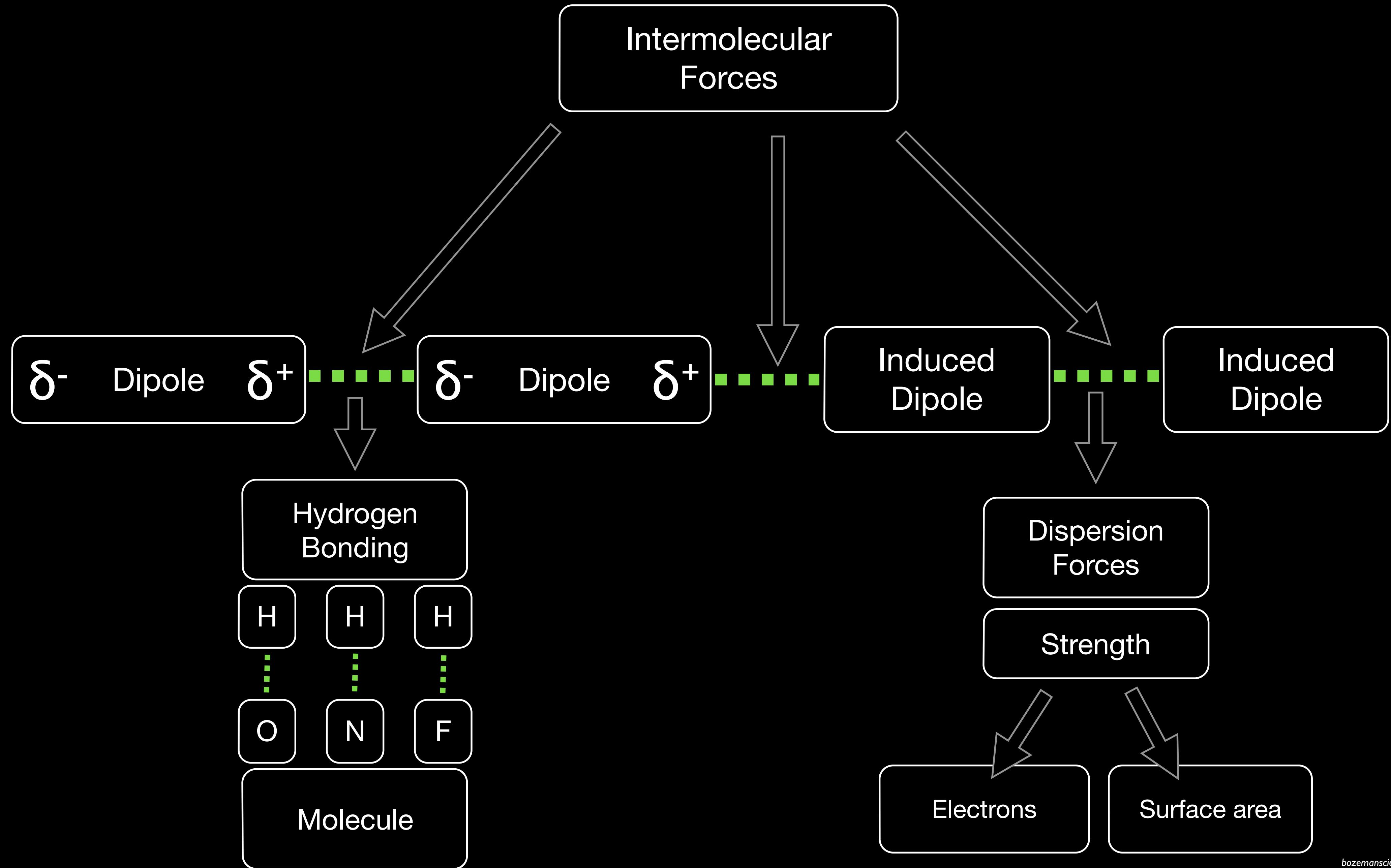
Intermolecular Potential Energy



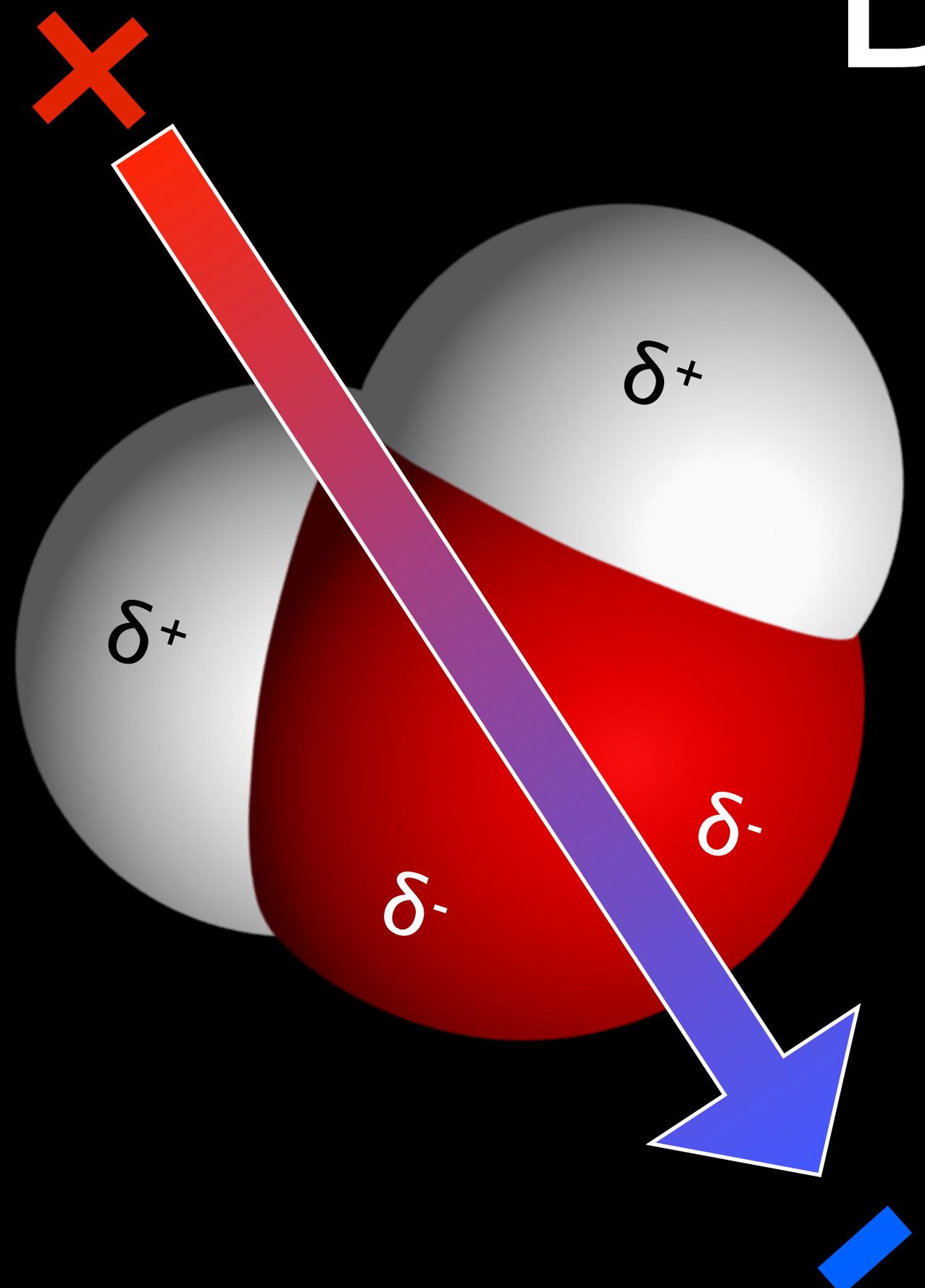
Chemistry Essentials - 054

Intermolecular Forces

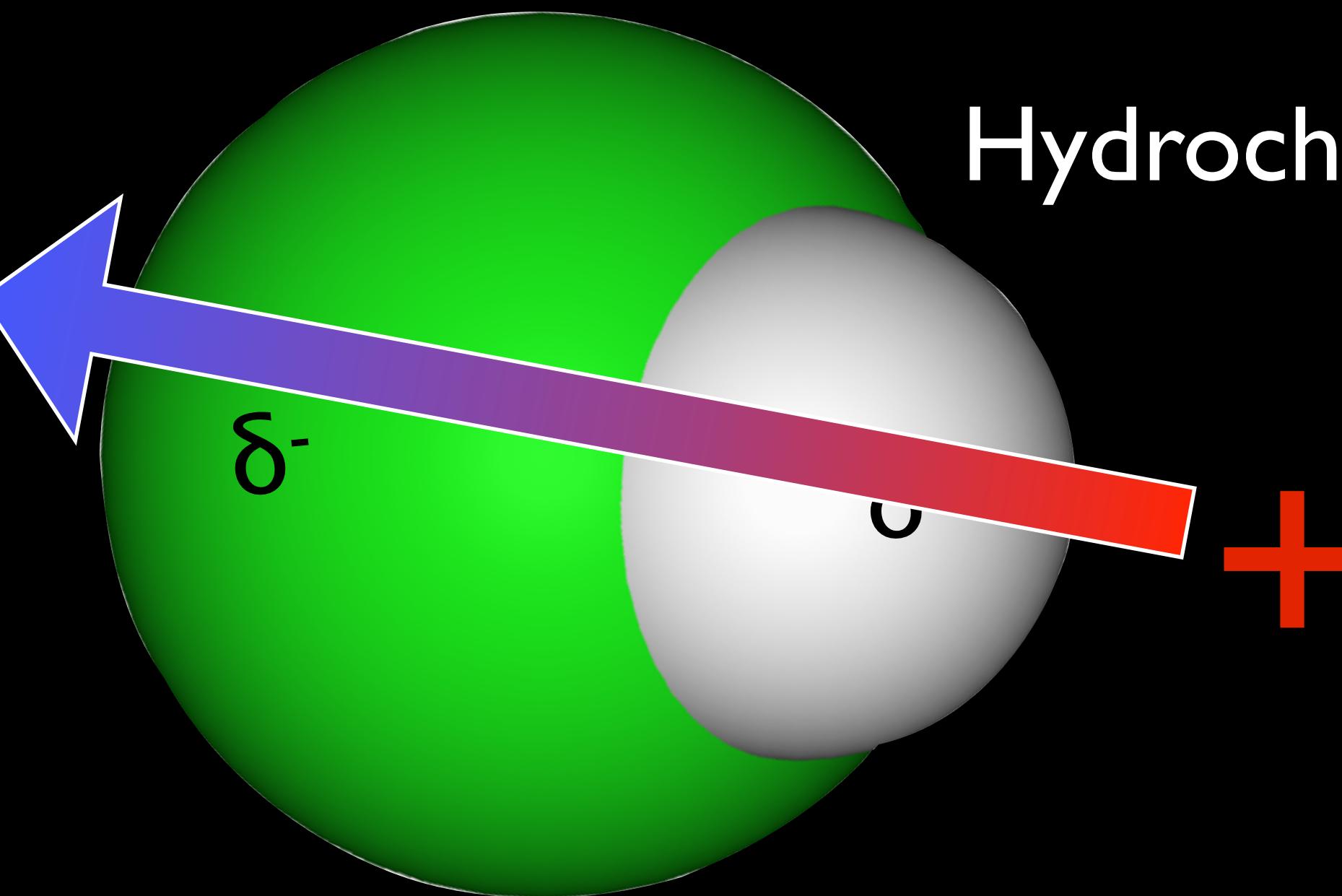




Dipole

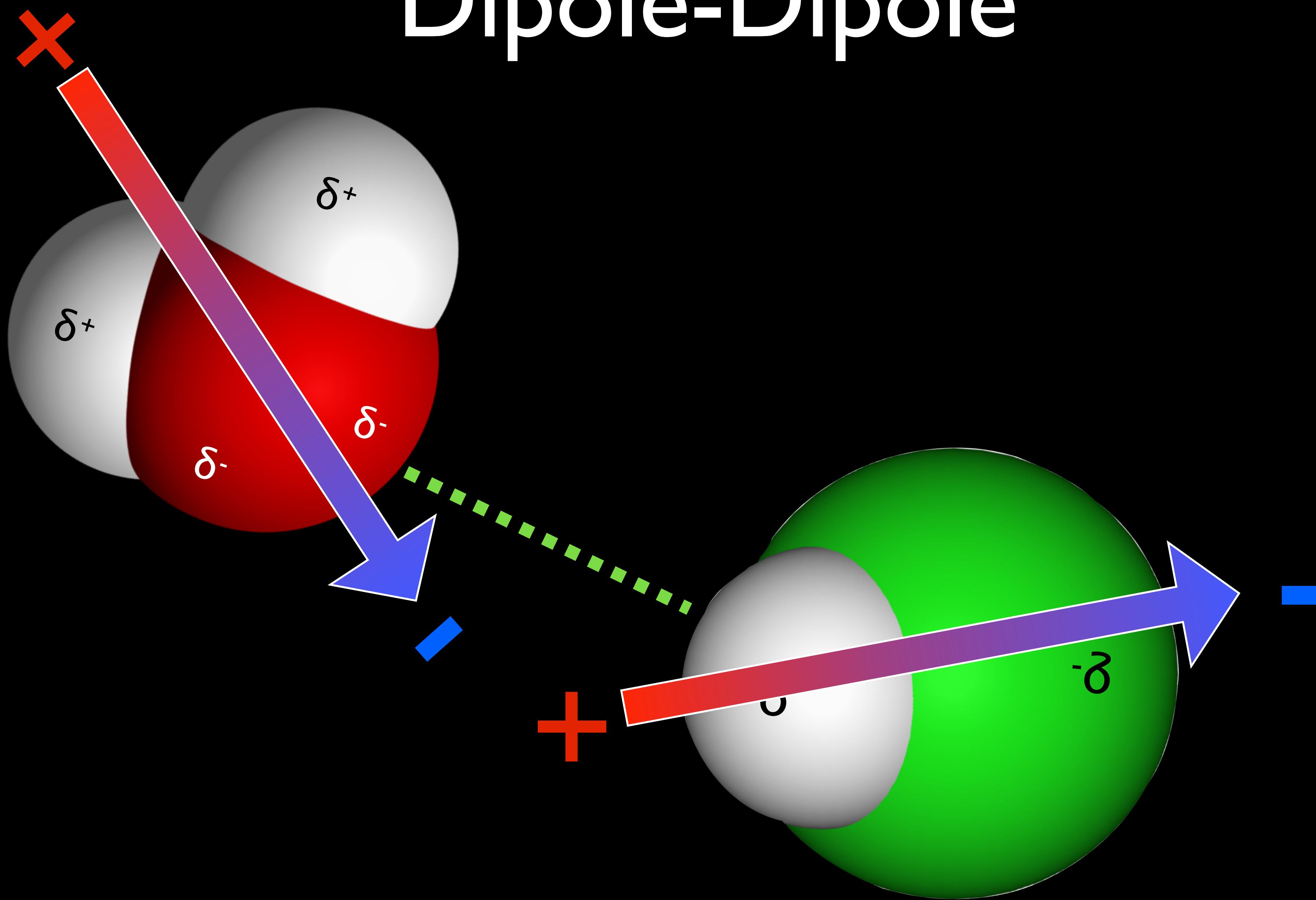


Water

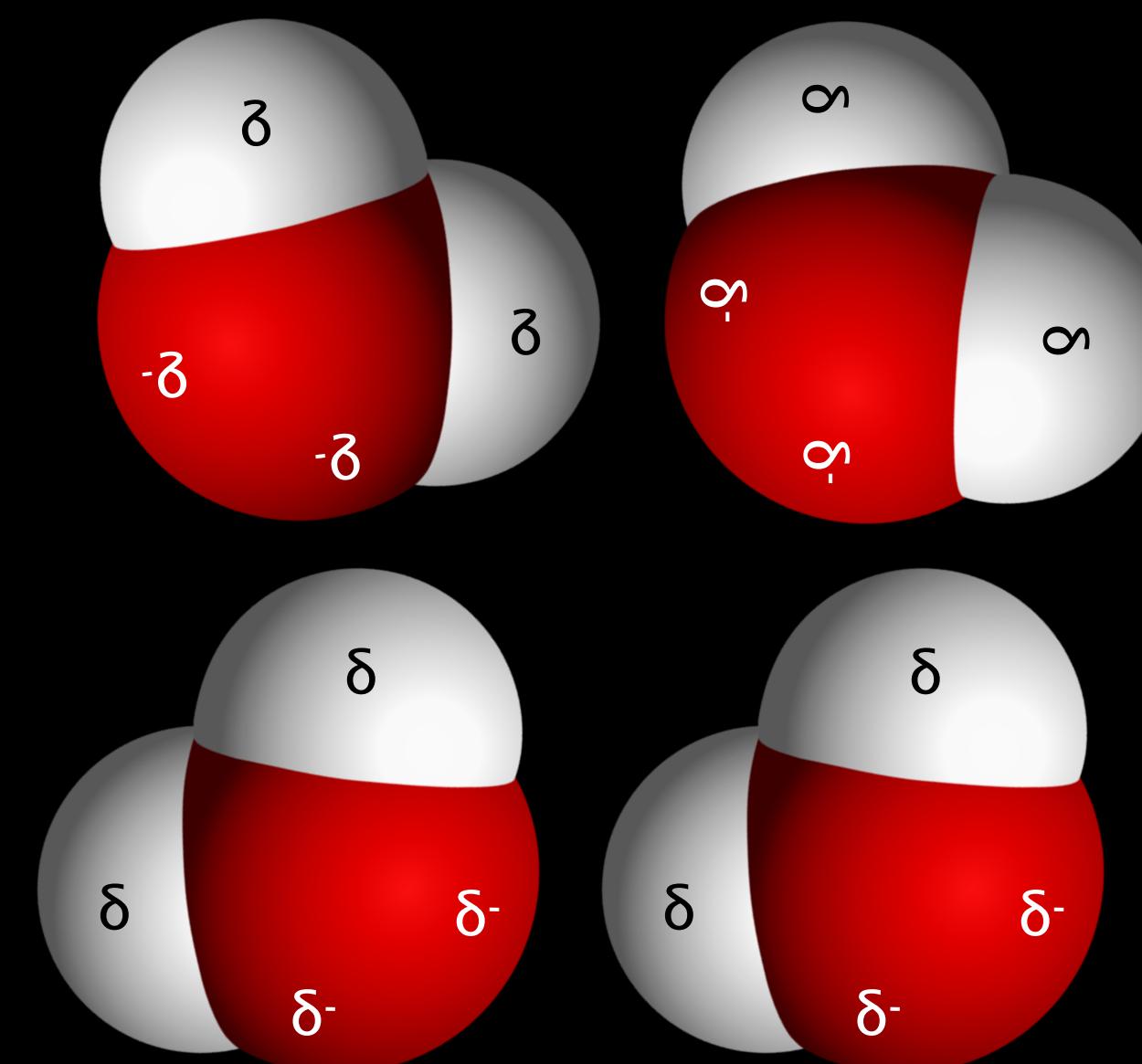
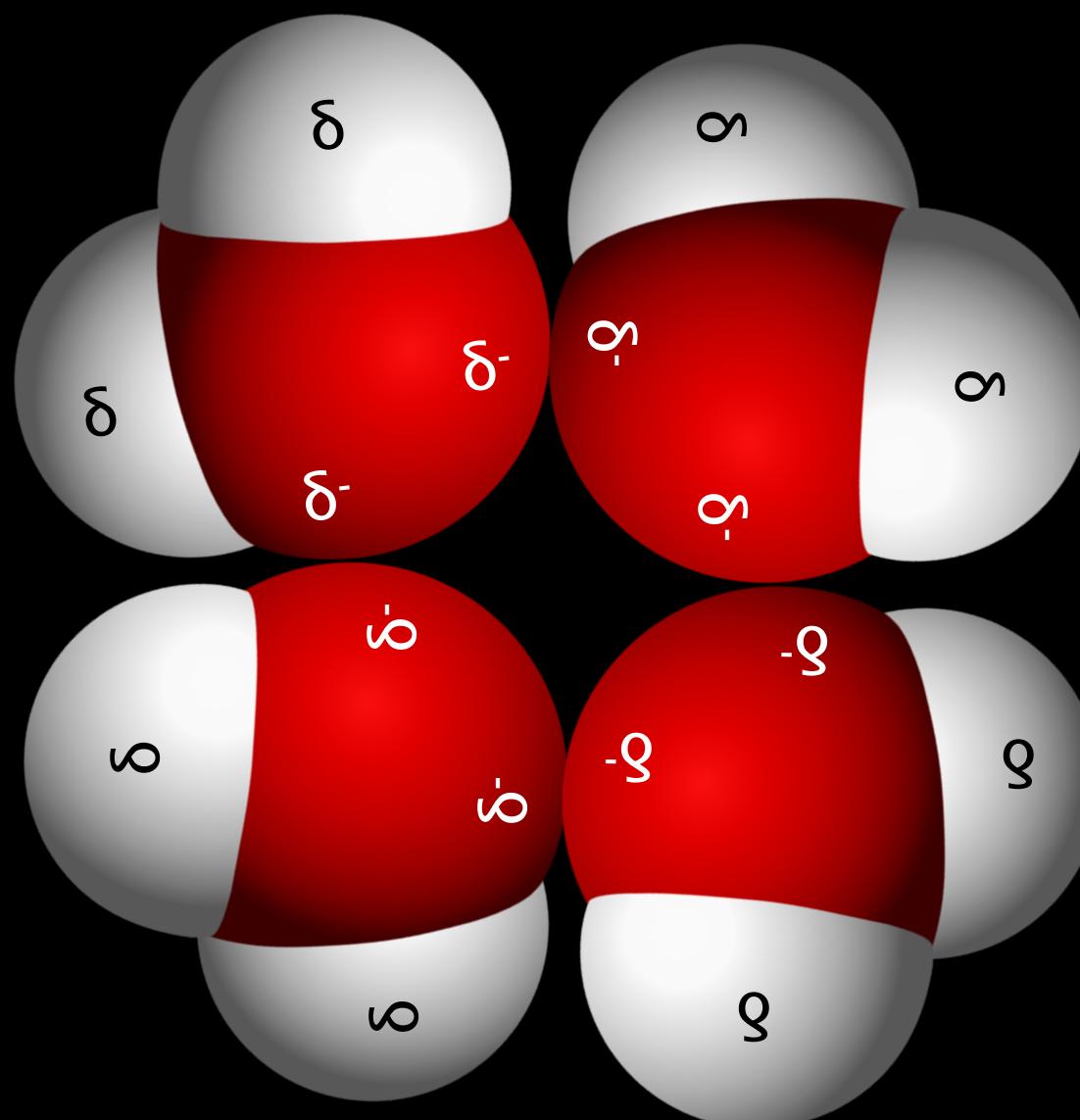


Hydrochloric Acid

Dipole-Dipole



Hydrogen Bonds



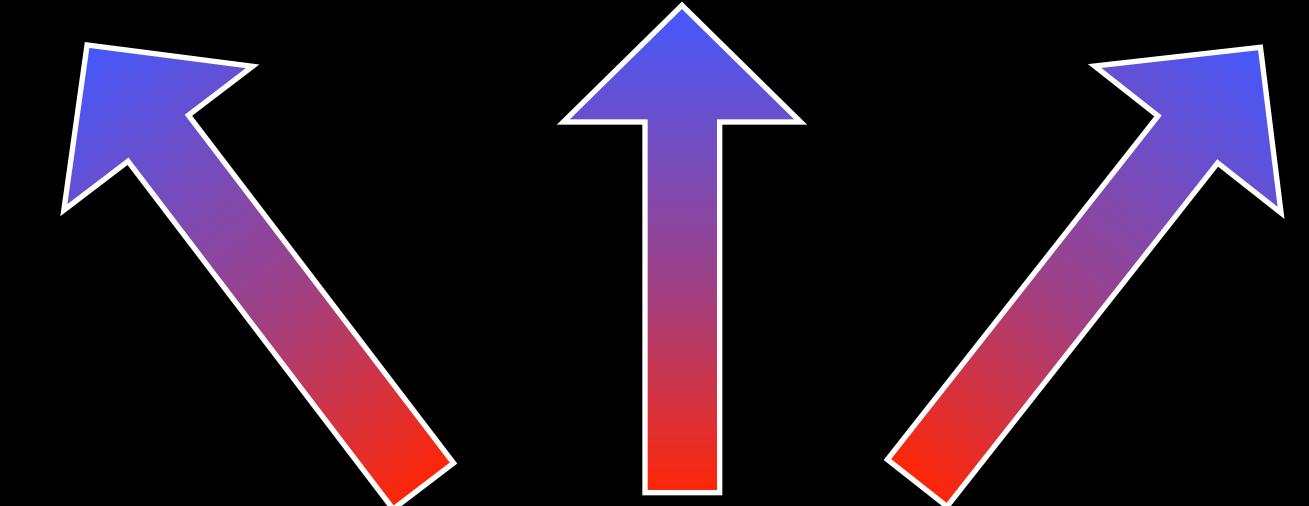
δ^-

7	N
Nitrogen	14.01

8	O
Oxygen	16.00

9	F
Fluorine	19.00

δ^-

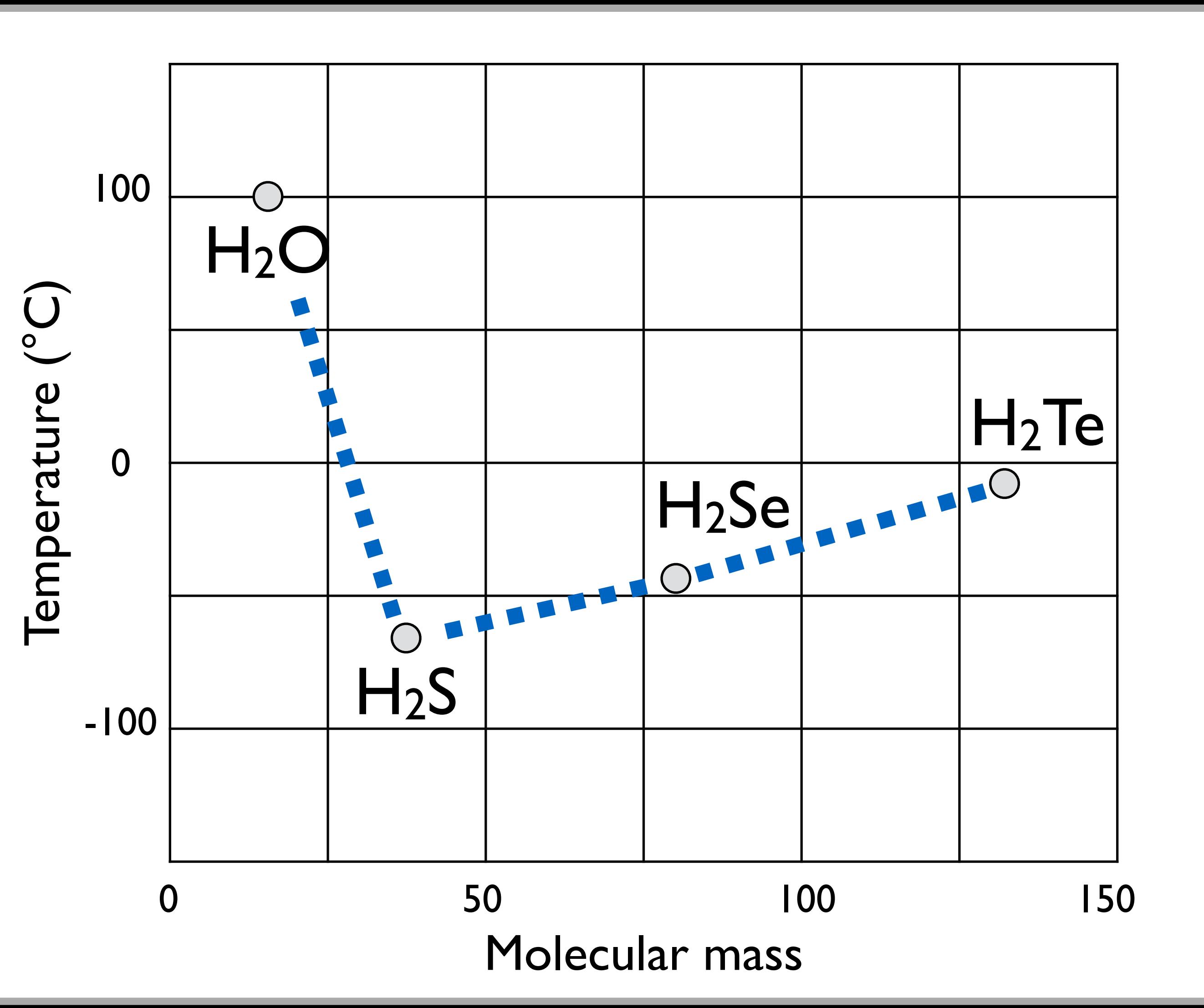


δ^+

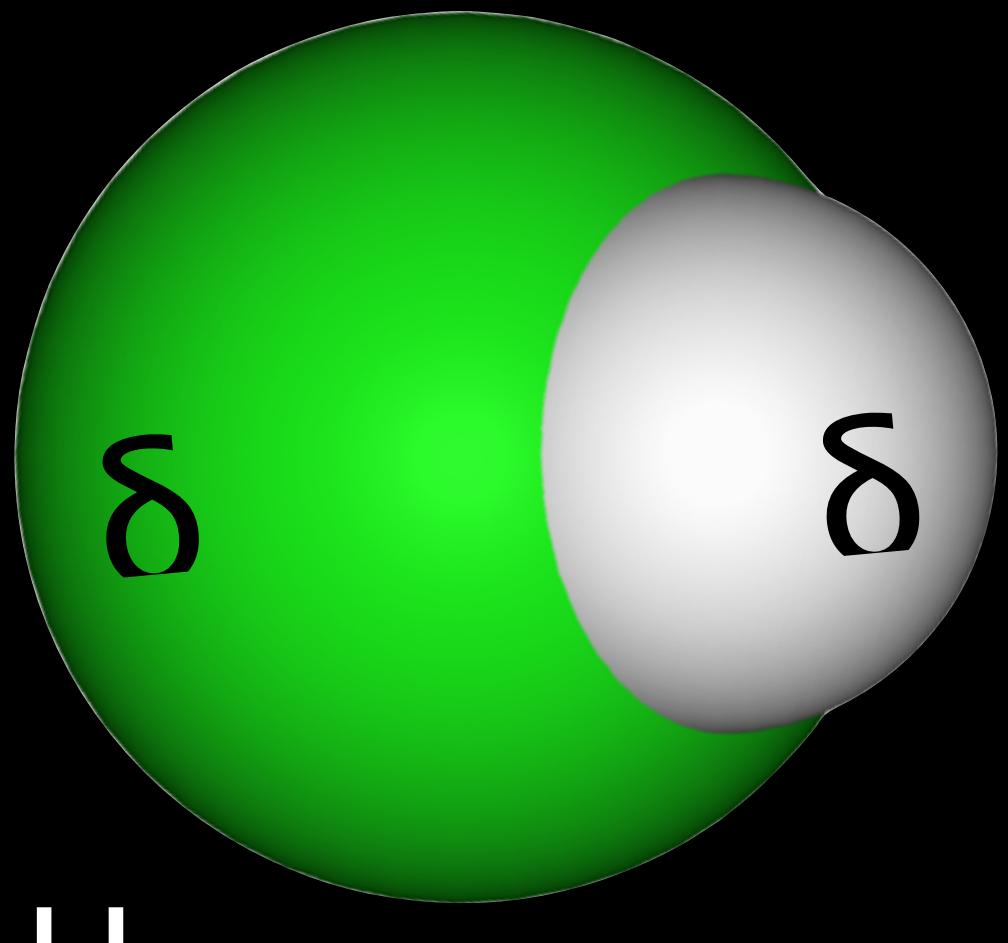
1	H
Hydrogen	1.01

Boiling Point

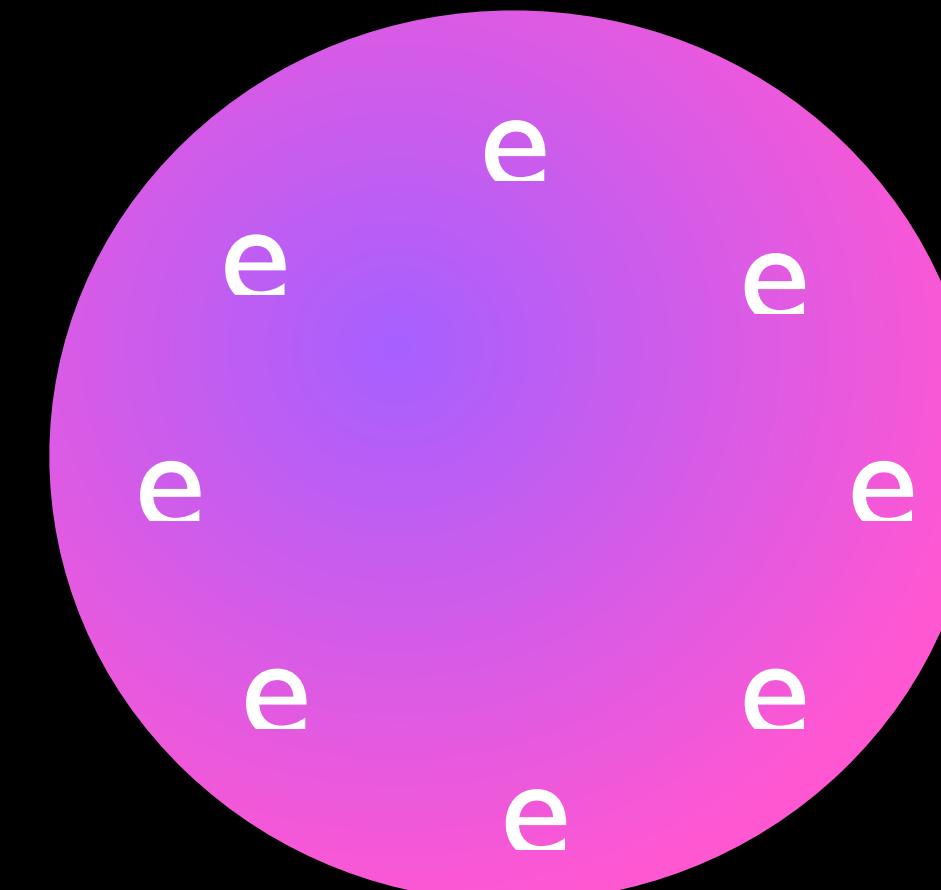
H	1 (1.01)
O	8 (15.99)
S	16 (32.06)
Se	34 (78.96)
Te	52 (127.60)



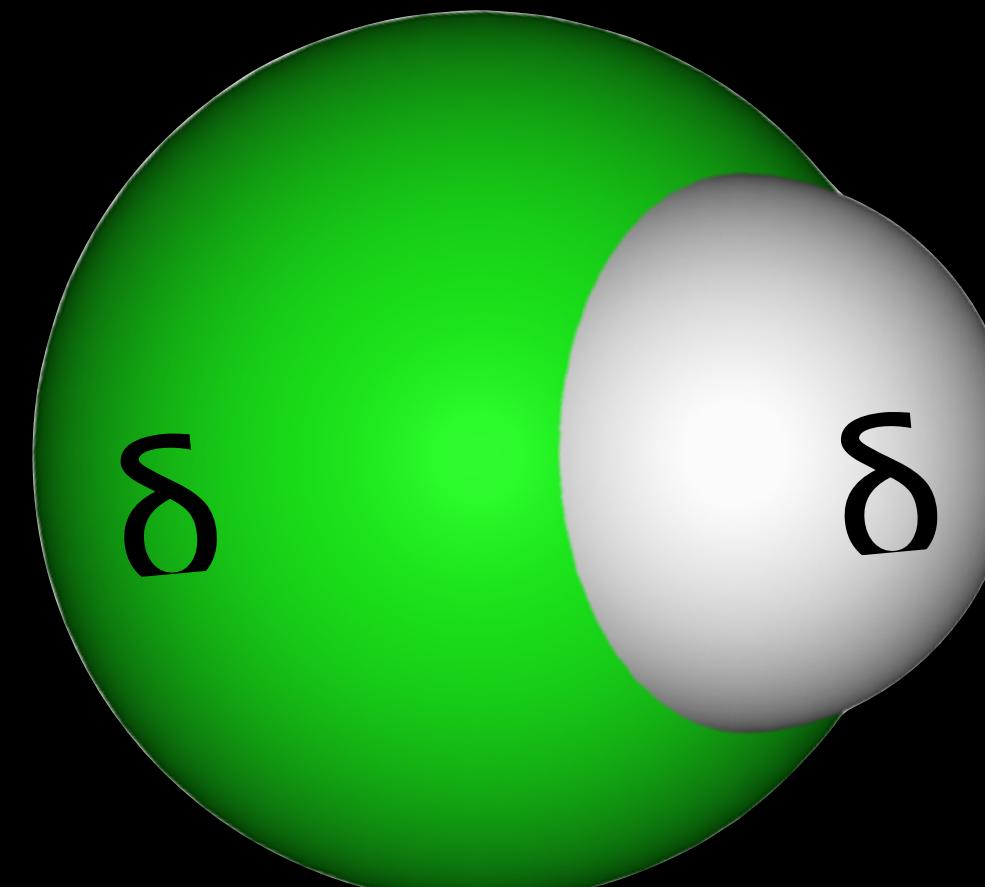
Dipole-Induced dipole force



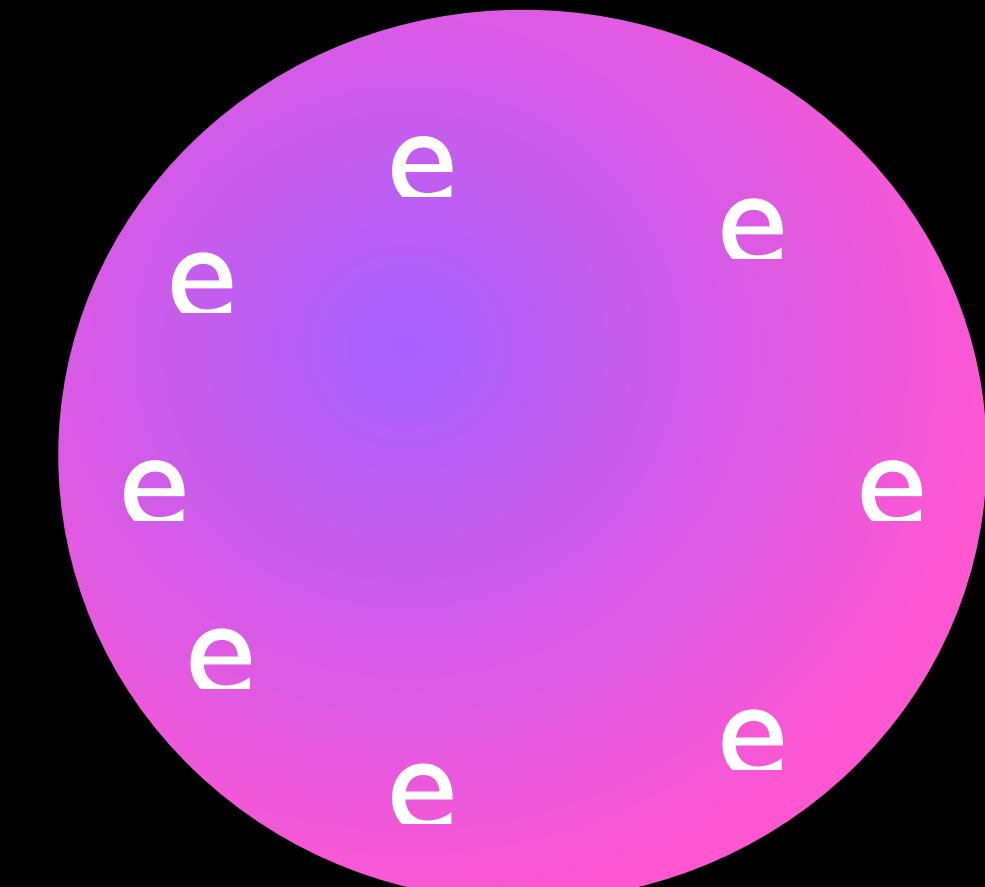
H



A

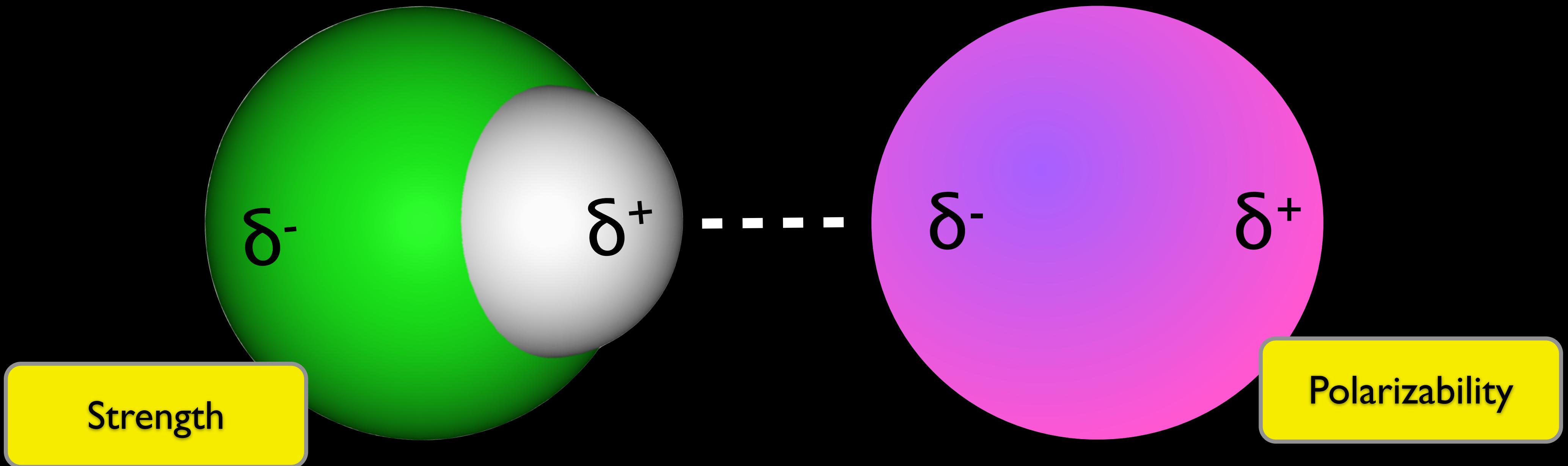


H

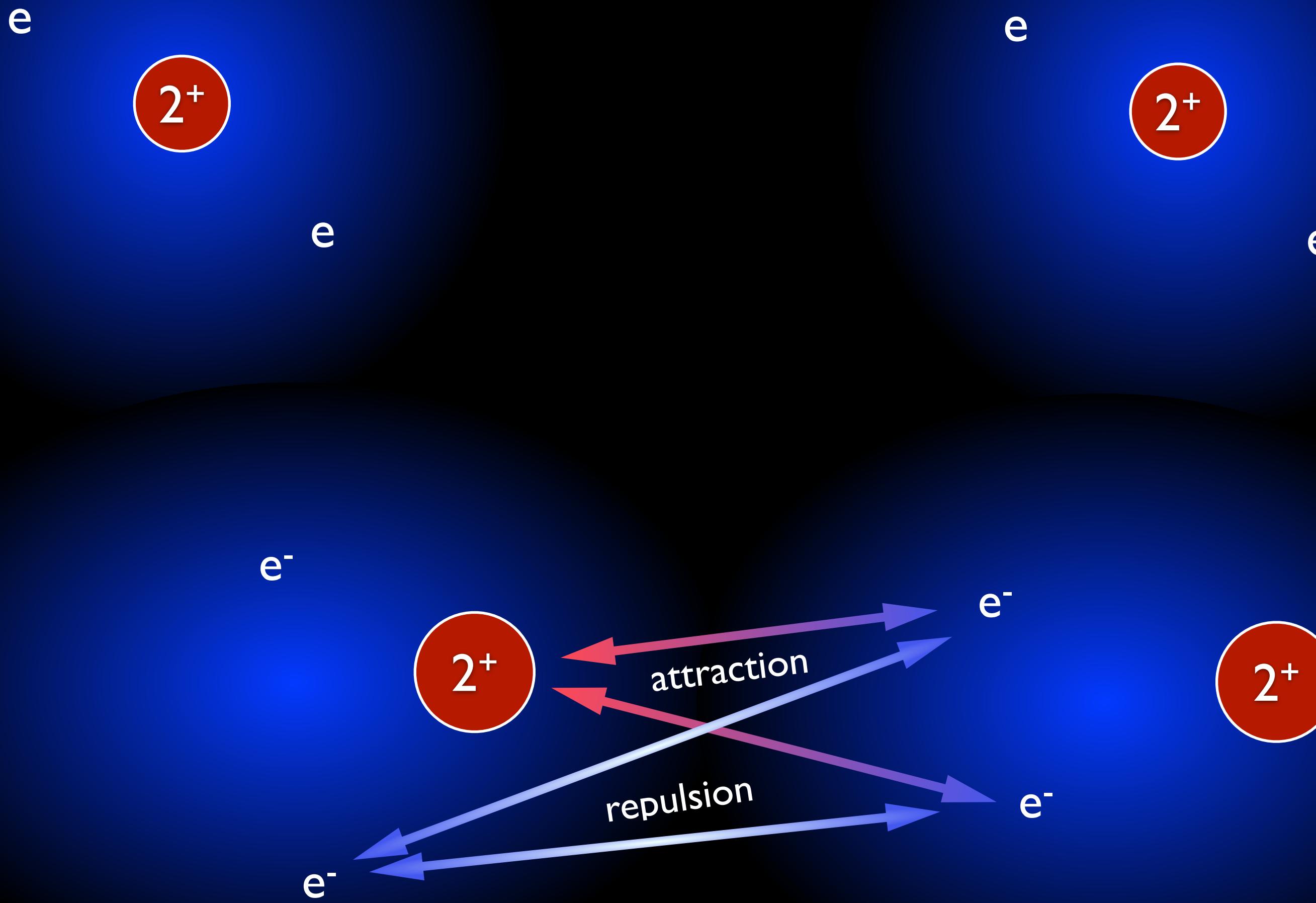


A

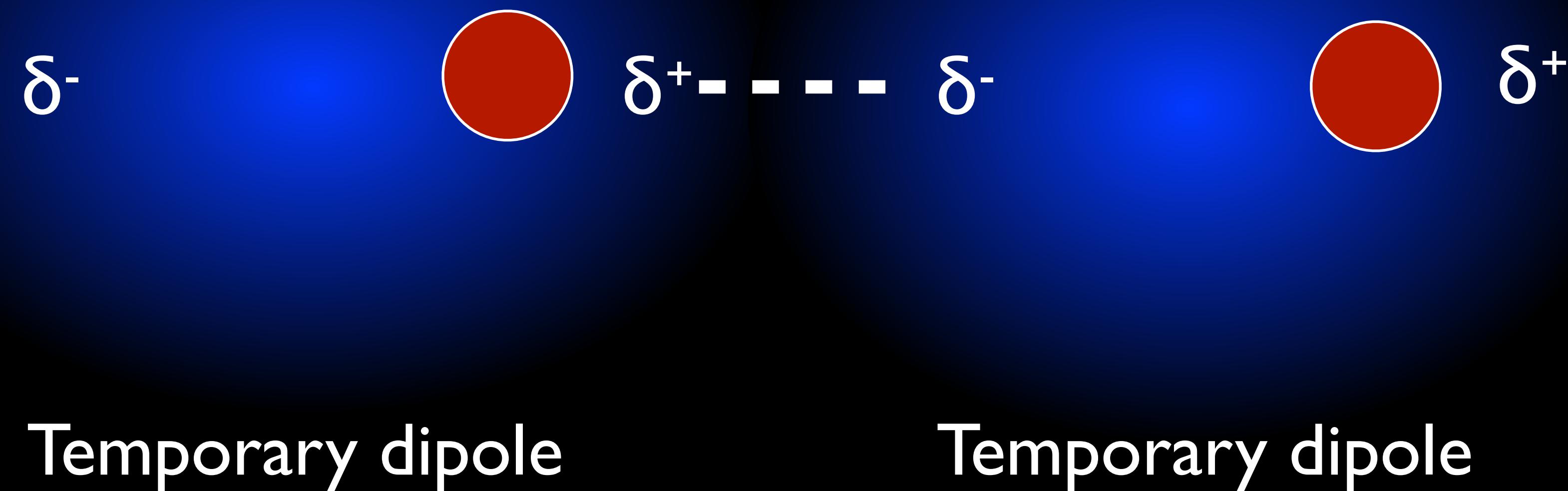
Dipole-Induced dipole force



Induced dipole-Induced dipole force



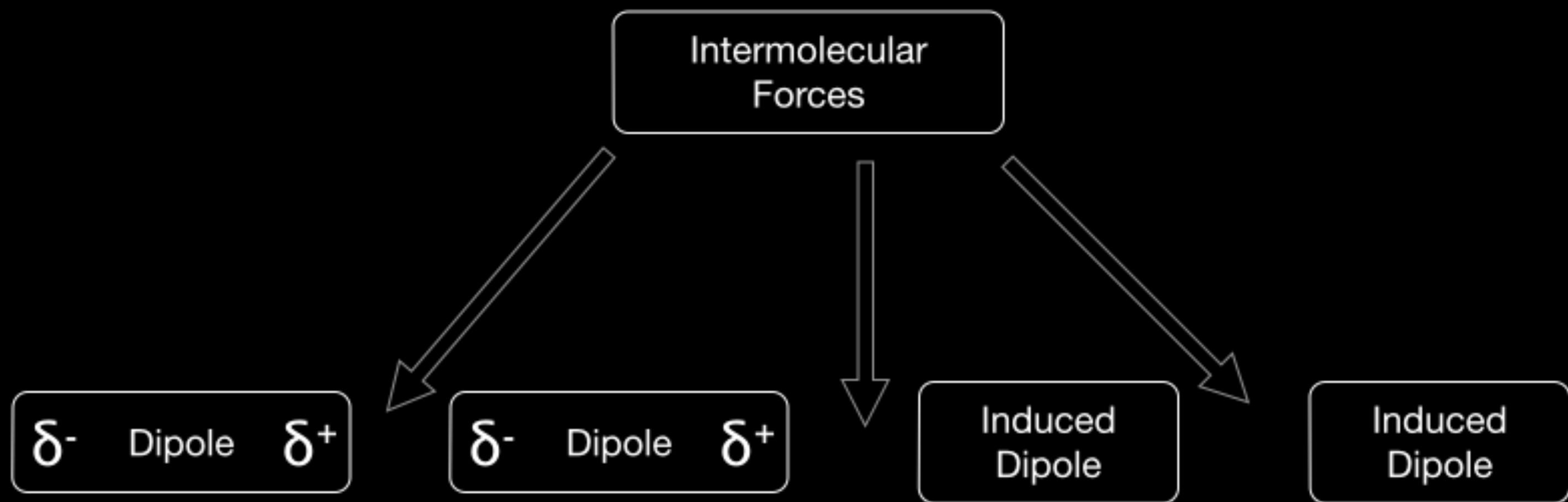
Induced dipole-Induced dipole force



- Electrons
- Surface area



Did you learn?



To make claims or predictions regarding intermolecular forces based on type of molecule and number of electrons.

Acknowledgements

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Dbc334. *English: Molecule of Water.*, November 2, 2006. File:Water-3D-vdW.png. http://commons.wikimedia.org/wiki/File:Water_molecule_3D.svg.
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“File:Uroplatus Fimbriatus (3).jpg,” December 22, 2013. [http://en.wikipedia.org/wiki/File:Uroplatus_fimbriatus_\(3\).jpg](http://en.wikipedia.org/wiki/File:Uroplatus_fimbriatus_(3).jpg).



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